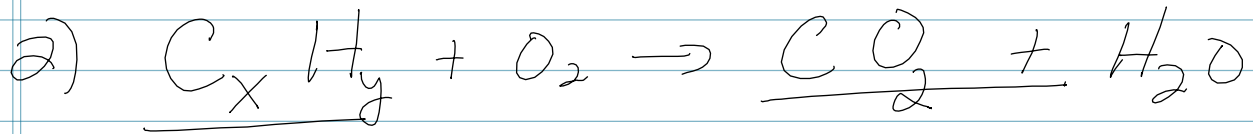
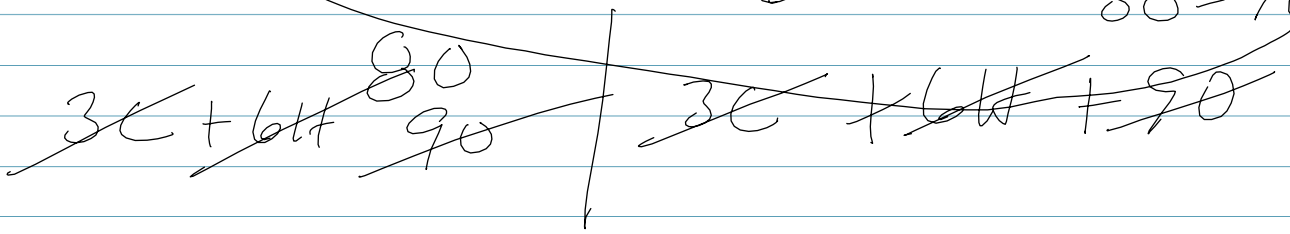
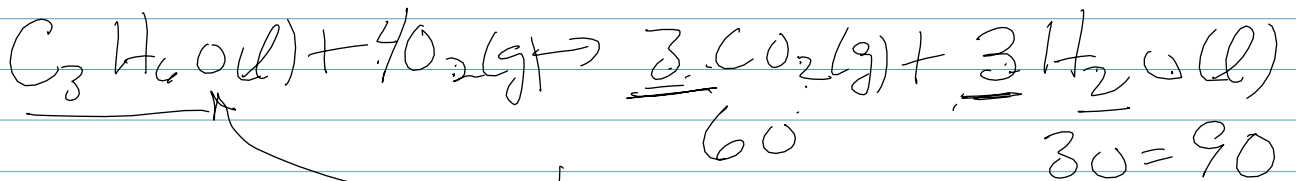
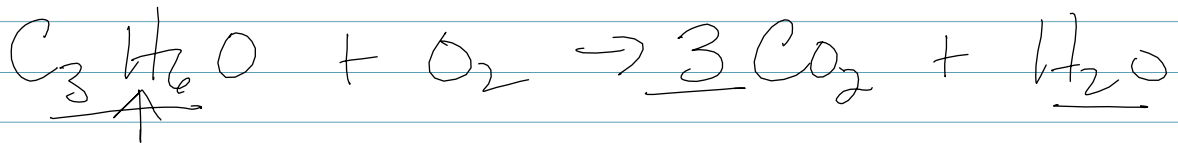
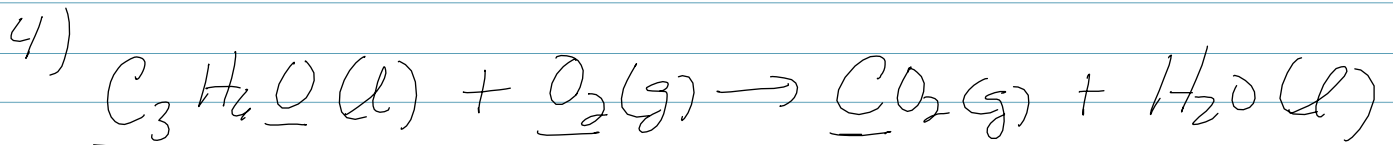
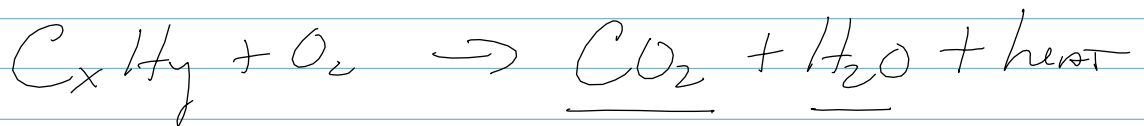


Hydrocarbon — Carbon + Hydrogen
(sometimes O)



3.) Extra Heat



5. 5.84g C₅H₁₂ releases 286 kJ

$$\frac{286 \text{ kJ}}{5.84 \text{ g}} \times \frac{\text{g}}{\text{mol}} \cdot C_5H_{12} \cdot 1.011 \frac{\text{g}}{\text{mol}}$$

$$(5 \times 12.01 \frac{\text{g}}{\text{mol}}) + (12 \times 1.01 \frac{\text{g}}{\text{mol}}) = 72.06 \frac{\text{g}}{\text{mol}}$$